



CHEMISTRY

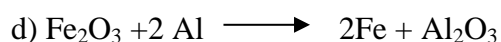
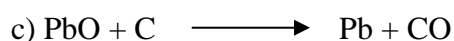
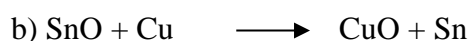
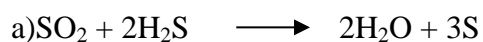
CH: 1 CHEMICAL REACTIONS AND EQUATIONS

Name: _____

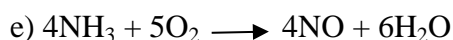
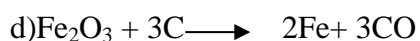
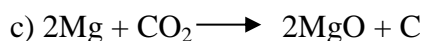
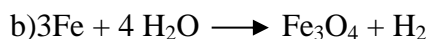
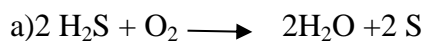
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1. Identify the substance oxidized and reduced in the following reactions.



2. Write down the oxidizing and reducing agents in the following reactions.



3. Name the following.

a) the gas evolved when lead nitrate is heated.

a) the reaction that occurs when food becomes rancid.

b) the chemicals used in black and white photography.

c) the black substance obtained when sugar is heated in a dry test tube.

4. What happens when

a) FeSO_4 is heated?

b) quick lime is added to water?



INDIAN SCHOOL NIZWA - WORKSHEET

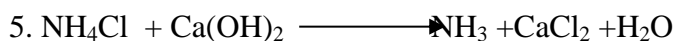
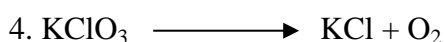
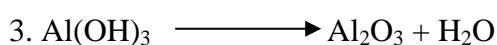
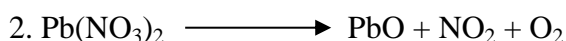
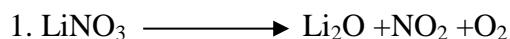
c) silver chloride is exposed to air?

5. Give reasons.

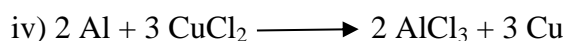
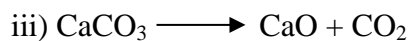
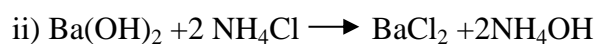
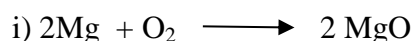
a) Silver articles become black after some time when exposed to air.

b) Copper sulphate solution turns colourless when zinc metal is dipped in it and kept for some time.

6. Balance the following equations



7. Identify the type of reaction.



8. Translate the following reaction into chemical equation, balance it and represent the Equation with state symbols.

a) Aluminium metal when added to sulphuric acid produces aqueous solution of aluminium sulphate and hydrogen gas.

b) Ammonia gas combines with oxygen gas to form nitrogen gas and liquid water.

9. Write the balanced equation and classify them as exothermic and endothermic reaction

a. When calcium oxide is added to water calcium hydroxide is produced and a rise in temperature is observed.

b. When barium hydroxide is mixed with ammonium chloride, barium chloride and ammonium hydroxide are produced and a fall in temperature is observed.

10. State one advantage and disadvantage of corrosion.